

ISO

INTERNATIONAL ORGANIZATION FOR STANDARDIZATION

ISO RECOMMENDATION R 981

SODIUM HYDROXIDE FOR INDUSTRIAL USE
DETERMINATION OF CHLORIDE CONTENT
VOLHARD VOLUMETRIC METHOD

1st EDITION

February 1969

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BRIEF HISTORY

The ISO Recommendation R 981, *Sodium hydroxide for industrial use – Determination of chloride content – Volhard volumetric method*, was drawn up by Technical Committee ISO/TC 47, *Chemistry*, the Secretariat of which is held by the Ente Nazionale Italiano di Unificazione (UNI).

Work on this question led, in 1966, to the adoption of a Draft ISO Recommendation.

In December 1966, this Draft ISO Recommendation (No. 1090) was circulated to all the ISO Member Bodies for enquiry. It was approved, subject to a few modifications of an editorial nature, by the following Member Bodies :

Austria	Ireland	South Africa, Rep. of
Belgium	Israel	Spain
Chile	Italy	Switzerland
Cuba	Japan	Thailand
Czechoslovakia	Korea, Dem. P. Rep. of	Turkey
France	Netherlands	U.A.R.
Germany	New Zealand	United Kingdom
Hungary	Poland	U.S.S.R.
India	Portugal	Yugoslavia
Iran	Romania	

One Member Body opposed the approval of the Draft :

U.S.A.

The Draft ISO Recommendation was then submitted by correspondence to the ISO Council, which decided, in February 1969, to accept it as an ISO RECOMMENDATION.

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SODIUM HYDROXIDE FOR INDUSTRIAL USE

DETERMINATION OF CHLORIDE CONTENT

VOLHARD VOLUMETRIC METHOD

1. SCOPE

This ISO Recommendation describes the Volhard volumetric method for the determination of chloride content in sodium hydroxide for industrial use.

2. FIELD OF APPLICATION

The method is applicable to the determination of chloride content of sodium hydroxide for industrial use, for contents greater than 0.05 % (m/m), expressed as sodium chloride and calculated on NaOH.

Special case

Determination of chloride contents between 0.01 and 0.05 % (m/m), expressed as sodium chloride and calculated on NaOH.

3. PRINCIPLE

Precipitation of Cl^- ions by addition of excess silver nitrate and titration of the excess with ammonium thiocyanate in the presence of ammonium-iron (III) sulphate.

4. REAGENTS

Distilled water or water of equivalent purity should be used in the test.

4.1 *Nitric acid*, approximately $d = 1.4$, 68 % (m/m) or 14 N solution.

4.2 *Nitric acid*, approximately 0.2 N solution.

4.3 *Silver nitrate*, 0.1 N standard volumetric solution (see Note in section 7).

4.4 *Ammonium thiocyanate*, 0.1 N standard volumetric solution (see Note in section 7).

4.5 *Ammonium-iron (III) sulphate*, nitric solution.

Dissolve 25 g of ammonium-iron (III) sulphate hydrate $[(\text{NH}_4)_2\text{SO}_4 \cdot \text{Fe}_2(\text{SO}_4)_3 \cdot 24\text{H}_2\text{O}]$ in 100 ml of water. Add 50 ml of nitric acid solution, approximately $d = 1.4$, and mix thoroughly.

4.6 *Litmus paper*.

5. APPARATUS

Ordinary laboratory apparatus.

6. PROCEDURE

6.1 Test portion

Transfer 100.0 ml of sample solution A* to a 500 ml conical flask.

6.2 Titration

Neutralize the sample solution (6.1) with the nitric acid solution (4.1) in the presence of litmus paper (4.6). Then add 0.5 to 1 ml excess nitric acid solution (4.1) and cool under running water to room temperature. Add approximately 3 ml of the ammonium-iron (III) sulphate solution (4.5).

Fill a burette with the silver nitrate standard volumetric solution (4.3) and another with the ammonium thiocyanate standard volumetric solution (4.4). Run 0.20 ml of the latter solution (4.4) into the sample solution so as to form ferric thiocyanate, a red compound the appearance of which indicates the end-point of the titration.

Then add the silver nitrate standard volumetric solution (4.3) until the reddish-pink colour disappears and add an excess of approximately 2 ml measured to the nearest 0.05 ml. Stir for 4 to 5 minutes and filter under reduced pressure on a glass filter crucible fitted with a sintered disk of porosity between 5 and 15 μm . Wash both the conical flask and the filter crucible three times with a small amount (approximately 10 ml) of the nitric acid solution (4.2) and then titrate the excess silver nitrate in the filtrate by adding the ammonium thiocyanate standard volumetric solution (4.4) until the pink colour reappears.

Deduct from the titration the drop that causes the end-point.

7. EXPRESSION OF RESULTS

Chloride content, expressed as sodium chloride (NaCl), is given as a percentage, by mass, by the following formula :

$$(V - V_1) \times A \times \frac{500}{100} \times \frac{100}{E} = 2.9221 \times \frac{(V - V_1)}{E}$$

where

V is the volume, in millilitres, of the silver nitrate standard volumetric solution (4.3) used;

V_1 is the volume, in millilitres, of the ammonium thiocyanate standard volumetric solution (4.4) used (0.20 ml + volume used for back-titration);

A is the mass, in grammes, of sodium chloride corresponding to 1 ml of 0.1 N silver nitrate solution (theoretical value 1 ml $\hat{=}$ 0.005 844 g of NaCl);

E is the mass, in grammes, of the test portion used for the preparation of the sample solution A*.

NOTE. — If the silver nitrate (4.3) and ammonium thiocyanate (4.4) standard volumetric solutions are not exactly the strength indicated in the list of reagents, suitable correction factors should be employed in calculating the result.

* See ISO Recommendation R 978, *Sodium hydroxide for industrial use — Preparation of sample solution*, clause 5.2.