



**International
Standard**

ISO 8044

**Corrosion of metals and alloys —
Vocabulary**

Corrosion des métaux et alliages — Vocabulaire

**Sixth edition
2024-12**

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Foreword

ISO (the International Organization for Standardization) is a worldwide federation of national standards bodies (ISO member bodies). The work of preparing International Standards is normally carried out through ISO technical committees. Each member body interested in a subject for which a technical committee has been established has the right to be represented on that committee. International organizations, governmental and non-governmental, in liaison with ISO, also take part in the work. ISO collaborates closely with the International Electrotechnical Commission (IEC) on all matters of electrotechnical standardization.

The procedures used to develop this document and those intended for its further maintenance are described in the ISO/IEC Directives, Part 1. In particular, the different approval criteria needed for the different types of ISO document should be noted. This document was drafted in accordance with the editorial rules of the ISO/IEC Directives, Part 2 (see www.iso.org/directives).

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For an explanation of the voluntary nature of standards, the meaning of ISO specific terms and expressions related to conformity assessment, as well as information about ISO's adherence to the World Trade Organization (WTO) principles in the Technical Barriers to Trade (TBT), see www.iso.org/iso/foreword.html.

This document was prepared by Technical Committee ISO/TC 156, *Corrosion of metals and alloys*, in collaboration with the European Committee for Standardization (CEN) Technical Committee CEN/TC 262, *Metallic and other inorganic coatings, including for corrosion protection and corrosion testing of metals and alloys*, in accordance with the Agreement on technical cooperation between ISO and CEN (Vienna Agreement).

This sixth edition cancels and replaces the fifth edition (ISO 8044:2020) which has been technically revised.

The main changes are as follows:

- several definitions have been changed, including the definition of corrosion;
- several editorial changes were made, including the addition of cross-references to other terms within the definitions.

Any feedback or questions on this document should be directed to the user's national standards body. A complete listing of these bodies can be found at www.iso.org/members.html.

Introduction

The definitions in this document have been drawn up with the objective of achieving a proper balance between precision and simplicity. The main objective of this document is to provide definitions that can be understood to have the same meaning by all concerned. Some corrosion terms in present use have developed through common usage and are not always logical. It has not, therefore, been possible to define certain terms in the form they are used in some countries. Because of the occasional conflicts between tradition and logic, some definitions inevitably represent a compromise.

An example of this kind of conflict is the term “corrosion”. This has been used to mean the process, the results of the process and the damage caused by the process. In this document, corrosion is understood to mean the process. Any detectable result of corrosion in any part of a corrosion system is termed “corrosion effect”. The term “corrosion damage” covers any impairment of the function of the technical system of which the metal and the environment form a part. Consequently, the term “corrosion protection” implies that the important thing is to avoid corrosion damage rather than to prevent corrosion, which in many cases is impossible and sometimes not necessary.

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Corrosion of metals and alloys — Vocabulary

1 Scope

This document defines terms relating to corrosion that are widely used in modern science and technology. In addition, some definitions are supplemented with short explanations.

Throughout the document, International Union of Pure and Applied Chemistry rules for electrode potential signs are applied. The term “metal” is also used to include alloys and other metallic materials.

Terms and definitions related to the inorganic surface treatment of metals are given in ISO 2080.

2 Normative references

There are no normative references in this document.

3 Terms related to corrosion in general

ISO and IEC maintain terminology databases for use in standardization at the following addresses:

- ISO Online browsing platform: available at <https://www.iso.org/obp>
- IEC Electropedia: available at <https://www.electropedia.org/>

3.1 corrosion

chemical or electrochemical interaction between a material, usually a metal, and its environment

Note 1 to entry: This can lead to deterioration of material properties, the environment or the function of the system containing that material.

3.2 corrosive agent

substance that initiates or promotes *corrosion* (3.1) when in contact with a given metal

3.3 corrosive environment

environment that contains one or more *corrosive agents* (3.2)

3.4 corrosion system

system consisting of one or more metals and those parts of the environment that influence *corrosion* (3.1)

Note 1 to entry: Parts of the environment can be, for example, coatings, surface layers or additional *electrodes* (7.1.2).

3.5 corrosion effect

change in any part of the *corrosion system* (3.4) caused by *corrosion* (3.1)

3.6 corrosion damage

corrosion effect (3.5) that causes impairment of the function of the metal, the environment or the technical system of which these form a part

3.7

corrosion failure

corrosion damage (3.6) characterized by the total loss of function of the technical system

3.8

corrosion product

substance formed as a result of *corrosion* (3.1)

3.9

scale

solid layer of *corrosion products* (3.8) formed on a metal at high temperature

Note 1 to entry: The term “scale” is also used in some countries for deposits from supersaturated water.

3.10

rust

visible *corrosion products* (3.8) consisting mainly of hydrated iron oxides

3.11

corrosion depth

distance between a point on the surface of a metal affected by *corrosion* (3.1) and the original surface of the metal

3.12

corrosion rate

corrosion effect (3.5) on a metal per unit time

Note 1 to entry: The unit used to express the corrosion rate depends on the technical system and on the type of corrosion effect. Thus, corrosion rate is typically expressed as an increase in *corrosion depth* (3.11) per unit time, or the mass of metal turned into *corrosion products* (3.8) per area of surface and per unit time, etc. The corrosion effect can vary with time and will not be the same at all points of the corroding surface. Therefore, reports of corrosion rates are typically accompanied by information on the type, time dependency and location of the corrosion effect.

3.13

corrosion resistance

ability of a metal to maintain *serviceability* (3.16) in a given *corrosion system* (3.4)

3.14

corrosivity

ability of an environment to cause *corrosion* (3.1) of a metal in a given *corrosion system* (3.4)

3.15

corrosion likelihood

qualitative and/or quantitative expression of the expected *corrosion effects* (3.5) in a defined *corrosion system* (3.4)

3.16

serviceability

ability of a *corrosion system* (3.4) to perform its specified functions without impairment due to *corrosion* (3.1)

3.17

durability

ability of a *corrosion system* (3.4) to maintain *serviceability* (3.16) over a specified time when the specified requirements for use and maintenance have been fulfilled

3.18

service life

time during which a *corrosion system* (3.4) meets the requirements for *serviceability* (3.16)

3.19

critical humidity

value of the relative humidity of an atmosphere above which there is a sharp increase in the *corrosion rate* (3.12) of a given metal

3.20

corrosion attack

corrosion effect (3.5) that is detrimental but has not progressed to the point of impairment of the function of the metal, the environment or the technical system of which these form a part

3.21

pickling

removal of oxides or other compounds from a metal surface by chemical or electrochemical action

3.22

pitting resistance equivalent number

PREN

indication of the resistance of stainless steels and nickel-based alloys to pitting in the presence of chloride-containing water

Note 1 to entry: An example formula for PREN is given by:

$$\text{PREN} = \% \text{Cr} + 3,3 [(\% \text{Mo}) + 0,5 (\% \text{W})] + 16 (\% \text{N})$$

Note 2 to entry: In general, the higher the PREN, the higher the resistance to *pitting corrosion* (4.15).

3.23

trap

micro structural site at which the residence time for a hydrogen atom is long compared to the residence time in an interstitial lattice site

3.24

time of wetness

period when a metallic surface is covered by adsorptive and/or liquid films of *electrolyte* (7.1.1) to be capable of causing *atmospheric corrosion* (4.4)

3.25

threshold stress

tensile stress above which stress corrosion cracks initiate and grow for specified test conditions

3.26

threshold stress intensity factor for stress corrosion cracking

K_{ISCC}

stress intensity factor above which stress corrosion crack propagation is sustained

Note 1 to entry: The threshold stress intensity factor is a concept of linear elastic fracture mechanics (LEFM) and is applicable when the plastic zone size is large compared with the microstructure, the crack is sufficiently long, and a high constraint to plastic deformation prevails, i.e. under plane strain predominant conditions. For growing stress corrosion cracks, LEFM is not necessarily applicable in all detail but is adopted as a pragmatic tool that is commonly used.

Note 2 to entry: Stress corrosion cracks can initiate at a surface or a surface defect and grow in the “small crack” regime at stress intensity factor levels below the apparent threshold stress intensity factor. Therefore, LEFM is not applicable in the “small crack” regime.

4 Terms related to types of corrosion

4.1

electrochemical corrosion

corrosion (3.1) involving at least one *anodic reaction* (7.1.9) and one *cathodic reaction* (7.1.6)

4.2

chemical corrosion

corrosion (3.1) not involving an electrochemical reaction

4.3

gaseous corrosion

corrosion (3.1) with dry gas as the *corrosive environment* (3.3) and without any liquid phase on the surface of the metal

4.4

atmospheric corrosion

corrosion (3.1) with the earth's atmosphere at ambient temperature as the *corrosive environment* (3.3)

4.5

marine corrosion

corrosion (3.1) with sea water as the main agent of the *corrosive environment* (3.3)

Note 1 to entry: This definition includes both immersed and splash zone conditions.

4.6

underground corrosion

corrosion (3.1) of buried metals, soil being the *corrosive environment* (3.3)

Note 1 to entry: The term "soil" includes not only the naturally occurring material but also any other material, such as ballast and backfill, used to cover a structure.

4.7

bacterial corrosion

microbiologically *influenced corrosion* (4.37) due to the action of bacteria

4.8

general corrosion

corrosion (3.1) proceeding over the whole surface of the metal exposed to the *corrosive environment* (3.3)

4.9

uniform corrosion

general corrosion (4.8) proceeding at almost the same rate over the whole surface

4.10

localized corrosion

corrosion (3.1) preferentially concentrated on discrete sites of the metal surface exposed to the *corrosive environment* (3.3)

Note 1 to entry: Localized corrosion can result in, for example, pits, cracks or grooves.

4.11

galvanic corrosion

corrosion (3.1) due to the action of a *corrosion cell* (7.1.13)

Note 1 to entry: The term has often been restricted to the action of bimetallic corrosion cells, i.e. to *bimetallic corrosion* (4.12).

4.12

bimetallic corrosion

DEPRECATED: contact corrosion

galvanic corrosion (4.11) where the *electrodes* (7.1.2) are formed by dissimilar metals

4.13

impressed current corrosion

electrochemical corrosion (4.1) due to the action of an external source of electric current

4.14

stray-current corrosion

impressed current corrosion (4.13) caused by current flowing through paths other than the intended circuits

4.15

pitting corrosion

localized corrosion (4.10) resulting in cavities extending from the surface into the metal

4.16

crevice corrosion

localized corrosion (4.10) associated with, and taking place in or immediately around, a narrow aperture or clearance formed between the metal surface and another surface (metallic or non-metallic)

4.17

deposit corrosion

localized corrosion (4.10) associated with, and taking place under or immediately around, a deposit of *corrosion products* (3.8) or other substance

4.18

water-line corrosion

corrosion (3.1) along, and as a consequence of the presence of, a gas/liquid boundary

4.19

selective corrosion

dealloying

corrosion (3.1) of an alloy whereby the components react in proportions that differ from their proportions in the alloy

4.20

dezincification of brass

selective corrosion (4.19) of brass resulting in the preferential removal of zinc

4.21

graphitic corrosion

selective corrosion (4.19) of grey cast iron resulting in the partial removal of metallic constituents and leaving graphite

4.22

intergranular corrosion

corrosion (3.1) in or adjacent to the grain boundaries of a metal

4.23

weld corrosion

corrosion (3.1) associated with the presence of a welded joint and taking place in the weld or its vicinity

4.24

knife-line corrosion

corrosion (3.1) resulting in a narrow slit in or adjacent to the filler/parent boundary of a welded or brazed joint

4.25

erosion corrosion

process involving conjoint *corrosion* (3.1) and erosion

Note 1 to entry: Erosion corrosion can occur in, for example, pipes with high fluid flow velocity and pumps and pipe lines carrying fluid containing abrasive particles in suspension or entrained in a gas flow.

4.26

cavitation corrosion

process involving conjoint *corrosion* (3.1) and cavitation

Note 1 to entry: Cavitation corrosion can occur, for example, in rotary pumps and on ships' propellers.

4.27

fretting corrosion

process involving conjoint *corrosion* (3.1) and oscillatory slip between two vibrating surfaces in contact

Note 1 to entry: Fretting corrosion can occur, for example, at mechanical joints in vibrating structures.

4.28

wear corrosion

process involving conjoint *corrosion* (3.1) and friction between two sliding surfaces in contact

4.29

corrosion fatigue

process involving conjoint *corrosion* (3.1) and alternating straining of the metal, often leading to cracking

Note 1 to entry: Corrosion fatigue can occur when a metal is subjected to cyclic straining in a *corrosive environment* (3.3).

4.30

stress corrosion

process involving conjoint straining of a material, usually a metal, due to applied or residual stress, and *corrosion* (3.1) of that material

4.31

stress corrosion cracking

cracking due to *stress corrosion* (4.30)

4.32

hydrogen embrittlement

process resulting in a decrease of the toughness or ductility of a metal due to absorption of hydrogen

Note 1 to entry: Hydrogen embrittlement often accompanies hydrogen formation, for example, by *corrosion* (3.1) or electrolysis. This can lead to cracking.

4.33

blistering

process resulting in a dome-shaped defect visible on the surface of an object and arising from localized loss of cohesion below the surface

Note 1 to entry: For example, blistering can occur on coated metal due to loss of adhesion between coating and substrate, caused by accumulation of products from *localized corrosion* (4.10). On uncoated metal, blistering can occur due to excessive internal hydrogen pressure.

4.34

spalling

fragmentation and detachment of portions of the surface layer or *scale* (3.9)

4.35

tarnishing

dulling, staining or discoloration of a metal surface, due to the formation of a thin layer of *corrosion products* (3.8)

4.36

aqueous corrosion

corrosion (3.1) with water or a water-based solution as the *corrosive environment* (3.3)

4.37

microbiologically influenced corrosion

MIC

corrosion (3.1) influenced by the action of microorganisms

Note 1 to entry: Compare with *bacterial corrosion* (4.7).

4.38

environmentally assisted cracking

cracking of a susceptible material, usually a metal, due to the conjoint action of its environment and mechanical stress

4.39

hydrogen-induced cracking

HIC

planar cracking that occurs in metals due to induced stresses when atomic hydrogen diffuses into the metal and then combines to form molecular hydrogen at *trap* (3.23) sites

4.40

hydrogen stress cracking

HSC

cracking that results from the presence of hydrogen in a metal and tensile stress (residual or applied or both)

Note 1 to entry: HSC describes cracking in metals that are not sensitive to *sulfide stress corrosion cracking* (4.43) but which can be embrittled by hydrogen when galvanically coupled, as the *cathode* (7.1.3), to another metal that is corroding actively as an *anode* (7.1.4). The term “galvanically induced HSC” has been used for this mechanism of cracking.

4.41

irradiation-assisted stress corrosion cracking

intergranular cracking of austenitic stainless steels resulting from a reduction in the chromium concentration in a very narrow band at the grain boundaries following exposure to high neutron irradiation doses exceeding one displacement per atom which causes the migration of point defects to the grain boundaries

4.42

stepwise cracking

SWC

cracking that connects *hydrogen-induced cracks* (HICs) (4.39) on adjacent planes in a metal

Note 1 to entry: This term describes the crack appearance. The linking of hydrogen-induced cracks to produce stepwise cracking is dependent upon local strain between the cracks and embrittlement of the surrounding metal by dissolved hydrogen. *HIC* (4.39)/*SWC* (4.42) is usually associated with low-strength plate steels used in the production of pipes and vessels.

4.43

sulfide stress corrosion cracking

SSCC

cracking of metal involving *corrosion* (3.1) and tensile stress, residual and/or applied, in the presence of water and hydrogen sulfide

Note 1 to entry: SSCC is a form of *hydrogen stress cracking* (4.40) and involves embrittlement of the metal by atomic hydrogen that is produced by acid corrosion on the metal surface. Hydrogen uptake is promoted in the presence of sulfides. The atomic hydrogen can diffuse into the metal, reduce ductility and increase susceptibility to cracking. High strength metallic materials and hard weld zones are prone to SSCC.

4.44

stress-oriented hydrogen-induced cracking

SOHIC

staggered small cracks formed approximately perpendicular to the principal stress (residual or applied) resulting in a “ladder-like” crack array linking (sometimes small) pre-existing *hydrogen-induced cracking* (HIC) (4.39) cracks

Note 1 to entry: The mode of cracking can be categorised as *stress corrosion cracking* (SCC) (4.31) caused by a combination of external stress and the local strain around hydrogen-induced cracks. SOHIC is related to SCC and HIC/*stepwise cracking* (4.42). It has been observed in parent material of longitudinally welded pipe and in the heat-affected zone of welds in pressure vessels. SOHIC is a relatively uncommon phenomenon usually associated with low-strength ferritic pipe and pressure vessel steels.

Note 2 to entry: Compare with *hydrogen embrittlement* (4.32).

[SOURCE: ISO 15156-1:2020, 3.23, modified — Note 2 to entry has been added.]

4.45

exfoliation corrosion

stratified form of subsurface *stress corrosion* (4.30) of susceptible primary wrought alloy mill products having a highly directional grain structure, accompanied by detachment of separate layers from the body of the material, formation of cracks and finally usually complete layer-by-layer disintegration of metal

Note 1 to entry: Exfoliation generally proceeds along grain boundaries, but with certain alloys and tempering it can develop along transgranular paths or a mixed intergranular/transgranular path.

Note 2 to entry: Layer corrosion can be developed on the first stage.

4.46

filiform corrosion

type of *corrosion* (3.1) proceeding under coating materials on metals in the form of threads, generally starting from bare edges or from local damage to the coating

Note 1 to entry: Usually, the threads are irregular in length and direction of growth, but they can also be nearly parallel and of approximately equal length. Filiform corrosion can occur under different *protective coatings* (5.5).

4.47

tribo-corrosion

any form of *corrosion* (3.1) that involves constant removal of the *passive layer* (7.3.6) due to fluid or particle impact on the corroding surface or the friction between the corroding surface and another surface

Note 1 to entry: Tribo-corrosion includes but is not restricted to: *wear corrosion* (4.28), *fretting corrosion* (4.27) and *erosion corrosion* (4.25).

Note 2 to entry: This process can result in an increase in friction of bearing surfaces in addition to causing material loss.

4.48

impingement attack

form of *erosion corrosion* (4.25) in aqueous liquids under high velocity or turbulent flow conditions on the metal surface causing repetitive disruption of protective films leading to accelerated *localised corrosion* (4.10)

4.49

high temperature corrosion

corrosion (3.1) by gases or deposits or both occurring at elevated temperatures under conditions where aqueous *electrolytes* (7.1.1) no longer exist

Note 1 to entry: High temperature corrosion can become significant at temperatures above 170 °C depending on material and environment.

4.50

hot corrosion

corrosion (3.1) by gases or deposits or both forming a liquid phase during *high temperature corrosion* (4.49) reaction

Note 1 to entry: Hot corrosion is a sub-term of high temperature corrosion.

Note 2 to entry: The most common liquid phases in which hot corrosion occurs are metal sulfates, metal vanadates and metal chlorides.

4.51

sulfidation

reaction of a metal or alloy with a sulfur-containing species to produce metal sulfides on or beneath the surface of the metal or alloy

4.52

metal dusting

carburation of metallic materials in process gases containing carbon oxides and hydrocarbons and with extremely low oxygen partial pressures leading to disintegration of the metal into dust of graphite, metal or carbides, or combinations

Note 1 to entry: The temperature range for metal dusting lies between 400 °C and 900 °C. For the mechanism to happen, a carbon activity higher than 1 in the process gas is required.

4.53

rebar corrosion

corrosion (3.1) of reinforcement bars in concrete

5 Terms related to corrosion protection

5.1

corrosion protection

modification of a *corrosion system* (3.4) so that *corrosion damage* (3.6) is reduced

5.2

degree of protection

reduction in *corrosion damage* (3.6) achieved by *corrosion protection* (5.1)

Note 1 to entry: All types of *corrosion* (3.1) present shall be considered.

Note 2 to entry: The degree of protection shall be expressed as a percentage.

5.3

temporary protection

corrosion protection (5.1) intended to last for a limited time only

Note 1 to entry: Temporary protection is used, for example, during storage and transportation of metal products or during shut-down of equipment.

5.4

protective layer

layer of a substance on a metal surface that decreases the *corrosion rate* (3.12)

Note 1 to entry: Such layers can be applied or arise spontaneously, for example, by *corrosion* (3.1).

5.5

protective coating

layer(s) of material applied to a metal surface to provide *corrosion protection* (5.1)

5.6

corrosion inhibitor

chemical substance that, when present in the *corrosion system* (3.4) at a suitable concentration, decreases the *corrosion rate* (3.12) without significantly changing the concentration of any *corrosive agent* (3.2)

Note 1 to entry: A corrosion inhibitor is generally effective in a small concentration.

5.7

volatile corrosion inhibitor

corrosion inhibitor (5.6) that can reach the metal surface in the form of vapour

5.8

de-aeration

removal of air from the environment

Note 1 to entry: If only oxygen is removed, the term “de-oxygenation” is more appropriate.

5.9

protective atmosphere

artificial atmosphere, of which the *corrosivity* (3.14) has been reduced by the removal or exclusion of *corrosive agents* (3.2) or by the addition of *corrosion inhibitors* (5.6)

5.10

critical potential

DEPRECATED: threshold potential

electrode potential (7.1.18) above or below which the *corrosion current* (7.2.3) varies markedly

Note 1 to entry: A critical potential is not an accurate value but can represent a small potential range of about 0,1 V width.

5.11

earth resistance

ohmic resistance between an *electrode* (7.1.2) in infinite space and a second, sufficiently remote electrode with infinite surface area, without consideration of *polarization resistance* (7.2.10)

Note 1 to entry: The earth resistance depends on the geometrical data of the first electrode and the resistivity of the medium within the space.

5.12

ground bed resistance

earth resistance (5.11) of a ground bed

Note 1 to entry: The ground bed is an *anode* (7.1.4) and its backfill is used for *cathodic protection* (7.4.3) in soil.

5.13

instant-off potential

electrode potential (7.1.18) that is measured immediately after switching off the *polarization* (7.1.28) current

Note 1 to entry: This is a *cathodic protection* (7.4.3) technique.

5.14

on-potential

non-current-times-resistance (IR)-free potential of the object to be protected, measured when the protection current is still flowing

Note 1 to entry: This is a *cathodic protection* (7.4.3) technique.

6 Terms related to corrosion testing

6.1

corrosion test

test carried out to assess the *corrosion resistance* (3.13) of a metal, the environmental contamination by *corrosion products* (3.8), the effectiveness of *corrosion protection* (5.1) or the *corrosivity* (3.14) of an environment

6.2

field corrosion test

corrosion test (6.1) conducted in a natural environment, such as air, water or soil

6.3

service corrosion test

corrosion test (6.1) conducted in service

6.4

simulated corrosion test

corrosion test (6.1) conducted under simulated service conditions

6.5

accelerated corrosion test

corrosion test (6.1) carried out under more severe conditions that will yield results in a shorter time than in service

6.6

cyclic atmospheric corrosion test

accelerated corrosion test (6.5) by exposure of coated and uncoated specimens to alternating environmental conditions designed to compare performance in atmospheric exposure

6.7

salt spray test

accelerated corrosion test (6.5) in a test cabinet by spraying the test specimens or objects with a solution containing sodium chloride

Note 1 to entry: There are three types of salt spray tests according to ISO 9227:

- neutral salt spray (NSS) test with (50 ± 5) g/l sodium chloride;
- acetic acid salt spray (AASS) test with the test solution for NSS test acidified with glacial acetic acid;
- copper-accelerated acetic acid salt spray (CASS) test with the test solution for AASS test completed with copper (II) chloride dihydrate ($\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$).

6.8

SCAB test

outdoor *accelerated corrosion test* (6.5) by intermittent spraying test specimens or objects with a spray of sodium chloride solution

Note 1 to entry: The method is suitable for specimens and objects of metals with or without coating.

Note 2 to entry: The term "SCAB" is an acronym of simulated corrosion atmospheric breakdown.

6.9

slow strain rate test

test for evaluating the susceptibility of a metal to *stress corrosion cracking* (4.31), which most commonly involves pulling a tensile specimen to failure in a representative environment at a constant displacement rate chosen to generate nominal strain rates usually in the range 10^{-5} s^{-1} to 10^{-8} s^{-1}

Note 1 to entry: Slow strain rate testing can also be applied to other specimen geometries, e.g. bend specimens.

7 Terms related to electrochemical matters

7.1 Electrochemical cell

7.1.1

electrolyte

medium in which an electric current is transported by ions

7.1.2

electrode

electronic conductor in contact with an *electrolyte* (7.1.1)

Note 1 to entry: In the electrochemical sense, the electrode is in fact restricted to narrow regions on both sides of the interface of this system.

7.1.3

cathode

electrode (7.1.2) at which a *cathodic reaction* (7.1.6) predominates

7.1.4

anode

electrode (7.1.2) at which an *anodic reaction* (7.1.9) predominates

7.1.5

electrode reaction

transfer of charge between an electronic conductor and an *electrolyte* (7.1.1)

7.1.6

cathodic reaction

transfer of negative charge from the electronic conductor to the *electrolyte* (7.1.1)

Note 1 to entry: Current enters the electronic conductor from the electrolyte. A cathodic reaction is a *reduction* (7.1.7) process, e.g. $\frac{1}{2} \text{O}_2 + \text{H}_2\text{O} + 2\text{e}^- \rightarrow 2\text{OH}^-$.

7.1.7

reduction

process in which a reactant accepts one or more electrons

7.1.8

reducing agent

substance causing *reduction* (7.1.7) by donating electrons

Note 1 to entry: During reduction, the reducing agent is oxidized.

7.1.9

anodic reaction

electrode reaction (7.1.5) equivalent to a transfer of positive charge from the electronic conductor to the *electrolyte* (7.1.1)

Note 1 to entry: Current enters the electrolyte from the electronic conductor. An anodic reaction is an *oxidation* (7.1.10) process. An example common in *corrosion* (3.1) is: $\text{M} \rightarrow \text{M}^{n+} + n\text{e}^-$.

7.1.10

oxidation

process in which a reactant loses one or more electrons

7.1.11

oxidizing agent

substance causing *oxidation* (7.1.10) by accepting electrons

Note 1 to entry: During oxidation, the oxidizing agent is reduced.

7.1.12

galvanic cell

combination of different *electrodes* (7.1.2) connected in series with an *electrolyte* (7.1.1)

Note 1 to entry: The galvanic cell is an electrochemical source of electrical current. It produces a current when the electrodes are connected by an external electronic conductor.

7.1.13

corrosion cell

short-circuited *galvanic cell* (7.1.12) in a *corrosion system* (3.4) where the corroding metal forms one of its *electrodes* (7.1.2)

Note 1 to entry: With respect to the distance between *anodes* (7.1.4) and *cathodes* (7.1.3):

- a) macrocells with dimensions from millimetres to several kilometres;
- b) microcells of microscopic dimensions can be distinguished.

7.1.14

concentration cell

corrosion cell (7.1.13) in which the potential difference arises from a difference in the concentration of *corrosive agent(s)* (3.2) near its *electrodes* (7.1.2)

7.1.15

differential aeration cell

corrosion cell (7.1.13) in which the potential difference arises from a difference in the concentration of oxygen near its *electrodes* (7.1.2)

Note 1 to entry: In some cases, the differential aeration cell can result in an *active-passive cell* (7.1.17).

7.1.16

bimetallic cell

corrosion cell (7.1.13) where the two *electrodes* (7.1.2) are formed by dissimilar metals

7.1.17

active-passive cell

corrosion cell (7.1.13) with one part of a metal surface in an *active state* (7.3.10) acting as an *anode* (7.1.4), and another part of the surface in a *passive state* (7.3.3) acting as a *cathode* (7.1.3)

7.1.18

electrode potential

voltage measured in the external circuit between an *electrode* (7.1.2) and a *reference electrode* (7.1.19) in contact with the same *electrolyte* (7.1.1)

7.1.19

reference electrode

electrode (7.1.2) having a stable and reproducible potential that is used as a reference in the measurement of *electrode potentials* (7.1.18)

7.1.20

corrosion potential

electrode potential (7.1.18) of a metal in a given *corrosion system* (3.4)

Note 1 to entry: The term is used whether or not there is a net (external) electrical current flowing to or from the metal surface under consideration.

7.1.21

free corrosion potential

open-circuit potential

corrosion potential (7.1.20) in the absence of a net (external) electrical current flowing to or from the metal surface

Note 1 to entry: See [Figure A.1](#).

7.1.22

galvanic series

list of metals ordered according to their *free corrosion potentials* (7.1.21) under specified conditions

Note 1 to entry: Other electronic conductors can also be included.

7.1.23

pitting initiation potential

lowest value of the *corrosion potential* (7.1.20) at which pit initiation is possible on a passive surface in a given *corrosive environment* (3.3)

7.1.24

activation polarization

change of the *electrode potential* (7.1.18) due to charge transfer

7.1.25

anodic polarization

change of the *electrode potential* (7.1.18) in the positive direction caused by current flow

Note 1 to entry: Compare with *electrode polarization* (7.2.9).

7.1.26

cathodic polarization

change of the *electrode potential* (7.1.18) in the negative direction caused by current flow

Note 1 to entry: Compare with *electrode polarization* (7.2.9).

7.1.27

concentration polarization

portion of the *electrode polarization* (7.2.9) of a cell that is produced by concentration changes resulting from the passage of current through the *electrolyte* (7.1.1)

Note 1 to entry: Compare with *concentration cell* (7.1.14).

7.1.28

polarization

difference between the *electrode potential* (7.1.18) and the *open-circuit potential* (7.1.21) of the *electrode* (7.1.2) in the environment under consideration which can, for example, be caused by the flow of current from an external source

7.1.29

resistance polarization

polarization (7.1.28) due to ohmic resistance in the *corrosion systems* (3.4)

7.1.30

overvoltage

voltage difference between the measured *electrode potential* (7.1.18) and the *equilibrium potential* (7.1.33)

7.1.31

anode potential

electrode potential (7.1.18) of the *anode* (7.1.4) of a *corrosion cell* (7.1.13) or *electrochemical cell* (7.1.44)

7.1.32

cathode potential

electrode potential (7.1.18) of the *cathode* (7.1.3) of a *corrosion cell* (7.1.13) or *electrochemical cell* (7.1.44)

7.1.33

equilibrium potential

electrode potential (7.1.18) of an *electrode* (7.1.2) that is in thermodynamic equilibrium with its environment

7.1.34

rest potential

corrosion potential (7.1.20) after time transients have subsided and near-equilibrium has been achieved

7.1.35

reference potential

potential difference between a *reference electrode* (7.1.19) and the standard hydrogen *electrode* (7.1.2)

Note 1 to entry: The reference potential of the standard hydrogen electrode is zero by convention.

7.1.36

redox potential

potential of a reversible oxidation-reduction reaction in a given *electrolyte* (7.1.1) recorded on a standard hydrogen electrode scale

Note 1 to entry:
$$\text{Ox} + ne^- \xrightleftharpoons[\text{oxidation}]{\text{reduction}} \text{Red}.$$

7.1.37

standard electrode potential

equilibrium potential (7.1.33) with all reactants at a unit activity ($a = 1$) and in the standard conditions ($T = 298 \text{ K}$, $P = 0,101\ 3 \text{ MPa}$)

7.1.38

stress corrosion cracking potential

potential below which *stress corrosion cracking* (4.31) does not take place

7.1.39

counter electrode

auxiliary electrode

electrode (7.1.2) commonly used in applied *polarization* (7.1.28) to balance the current passing to the *working electrode* (7.1.45)

Note 1 to entry: It is usually made from a non-corroding material.

7.1.40

half cell

electrode (7.1.2) on which a theoretical single *oxidation* (7.1.10) or *reduction* (7.1.7) half reaction occurs

Note 1 to entry: Two half cells connected form an *electrochemical cell* (7.1.44).

7.1.41

mixed electrode

electrode (7.1.2) where diverse *electrode reactions* (7.1.5) take place

7.1.42

Tafel slope

slope of the straight-line portion of an electrochemical *polarization curve* (7.2.8) (plotted in terms of logarithm of the *current density* (7.2.6) versus potential) corresponding to an activation-controlled reaction

Note 1 to entry: This is commonly expressed as volts per decade.

7.1.43

Luggin probe

Haber-Luggin capillary

Luggin-Haber capillary

small tube or capillary filled with *electrolyte* (7.1.1), terminating close to the metal surface under study and used to minimize potential drop between the *reference electrode* (7.1.19) and the test *electrode* (7.1.2)

7.1.44

electrochemical cell

system composed of an *anode* (7.1.4), a *cathode* (7.1.3) and an *electrolyte* (7.1.1)

7.1.45

working electrode

test *electrode* (7.1.2) in an *electrochemical cell* (7.1.44), designed for electrochemical tests

7.2 Reaction rates

7.2.1

anodic partial current

sum of all the currents corresponding to *anodic reactions* (7.1.9) on an *electrode* (7.1.2)

Note 1 to entry: See [Figure A.1](#).

7.2.2

cathodic partial current

sum of all the currents corresponding to electrochemical *cathodic reactions* (7.1.6) on the *electrode* (7.1.2)

Note 1 to entry: See [Figure A.1](#).

7.2.3**corrosion current**

anodic partial current ([7.2.1](#)) due to *metal oxidation* ([7.1.10](#))

Note 1 to entry: The *corrosion current density* ([7.2.6](#)) is equivalent to the rate of *electrochemical corrosion* ([4.1](#)) according to Faraday's law.

7.2.4**free corrosion current**

corrosion current ([7.2.3](#)) at the *free corrosion potential* ([7.1.21](#))

Note 1 to entry: See [Figure A.1](#).

7.2.5**net current**

sum of all the *anodic partial current* ([7.2.1](#)) and *cathodic partial currents* ([7.2.2](#)) on an *electrode* ([7.1.2](#))

Note 1 to entry: The net current is dependent on many parameters, such as the *electrode potential* ([7.1.18](#)). See [Figure A.1](#).

7.2.6**current density**

current per area of an *electrode* ([7.1.2](#))

7.2.7**limiting current**

maximum electric current allowed by the slowest non-electrochemical step of a given *electrode* ([7.1.2](#)) process

7.2.8**current density/potential curve****polarization curve**

diagram relating *electrode potential* ([7.1.18](#)) to *current density* ([7.2.6](#))

7.2.9**electrode polarization**

change of *electrode potential* ([7.1.18](#))

Note 1 to entry: Often the *free corrosion potential* ([7.1.21](#)) is used as a reference value. The change can be caused, for example, by the application of an external electrical current or by the addition of an *oxidizing agent* ([7.1.11](#)) or *reducing agent* ([7.1.8](#)).

7.2.10**polarization resistance**

quotient of *electrode potential* ([7.1.18](#)) increment and the corresponding current increment

Note 1 to entry: Usually, the polarization resistance is measured in the vicinity of the *free corrosion potential* ([7.1.21](#)) (linear polarization measurement). The polarization resistance measured in this way can be related to the *free corrosion current* ([7.2.4](#)).

7.2.11**diffusion layer**

electrolyte ([7.1.1](#)) layer at the *electrode* ([7.1.2](#)) surface with a different concentration of a given species than that in the bulk of the solution

Note 1 to entry: In this layer, diffusion is an important transport mode for species formed or consumed at the electrode.

7.2.12**cathodic control**

limitation of the *corrosion rate* ([3.12](#)) by the rate of the *cathodic reaction* ([7.1.6](#))

7.2.13**anodic control**

limitation of the *corrosion rate* ([3.12](#)) by the rate of the *anodic reaction* ([7.1.9](#))

7.2.14

resistance control

limitation of the *corrosion rate* (3.12) by the ohmic resistance within the *corrosion cell* (7.1.13)

7.2.15

diffusion control

limitation of the *corrosion rate* (3.12) by the diffusion rate of *corrosive agents* (3.2) to, or *corrosion products* (3.8) from, the metal surface

7.2.16

mixed control

limitation of a *corrosion rate* (3.12) by the simultaneous action of two or more controlling factors

7.2.17

cell current

current that flows between an *anode* (7.1.4) and a *cathode* (7.1.3)

7.2.18

exchange current density

current density (7.2.6) at a single *electrode* (7.1.2) corresponding to the rate of internal charge transfer exchange within *anodic reactions* (7.1.9) and *cathodic reactions* (7.1.6) at their *equilibrium potential* (7.1.33)

7.3 Passivation

7.3.1

passivation

decrease of *corrosion rate* (3.12) by a *passivation layer* (7.3.6)

Note 1 to entry: Incomplete passivation can lead to *localized corrosion* (4.10).

7.3.2

passivator

chemical agent causing *passivation* (7.3.1)

7.3.3

passive state

passivity

state of a metal resulting from its *passivation* (7.3.1)

Note 1 to entry: See [Figure A.2](#).

7.3.4

passivation potential

corrosion potential (7.1.20) at which the *corrosion current* (7.2.3) has a peak value, and above which there is a range of potentials, where the metal is in a *passive state* (7.3.3)

Note 1 to entry: See [Figure A.2](#).

7.3.5

passivation current

corrosion current (7.2.3) at the *passivation potential* (7.3.4)

Note 1 to entry: If the *corrosive environment* (3.3) contains an *oxidizing agent* (7.1.11) being reduced at the passivation potential, the *net current* (7.2.5) measured will be lower than the passivation current.

7.3.6

passivation layer

passive layer

thin, adherent *protective layer* (5.4) formed on a metal surface through a reaction between the metal and its environment

7.3.7

depassivation

increase of *corrosion rate* (3.12) of a passive metal caused by general or local removal of its *passivation layer* (7.3.6)

7.3.8

re-activation

depassivation (7.3.7) caused by a decrease of the *electrode potential* (7.1.18)

7.3.9

depassivator

chemical agent causing *depassivation* (7.3.7)

7.3.10

active state

state of a corroding metal surface being below the *passivation potential* (7.3.4)

Note 1 to entry: See [Figure A.2](#).

7.3.11

re-activation potential

corrosion potential (7.1.20) below which *re-activation* (7.3.8) takes place

Note 1 to entry: See [Figure A.2](#).

7.3.12

transpassive state

state of a metal polarized to a potential value above the range of a *passive state* (7.3.3) characterized by a considerable increase of the *corrosion current* (7.2.3) and in the absence of *pitting corrosion* (4.15)

Note 1 to entry: See [Figure A.2](#).

7.3.13

transpassivation potential

corrosion potential (7.1.20) above which the metal is in a *transpassive state* (7.3.12)

Note 1 to entry: See [Figure A.2](#).

7.4 Electrochemical protection

7.4.1

electrochemical protection

corrosion protection (5.1) achieved by electrical control of the *corrosion potential* (7.1.20)

7.4.2

anodic protection

electrochemical protection (7.4.1) created by increasing the *electrode potential* (7.1.18) to a value within the potential range of the *passive state* (7.3.3)

7.4.3

cathodic protection

electrochemical protection (7.4.1) created by decreasing the *electrode potential* (7.1.18) to a level at which the *corrosion rate* (3.12) of the metal is significantly reduced

7.4.4

galvanic protection

electrochemical protection (7.4.1) in which the protection current is obtained from a *corrosion cell* (7.1.13) formed by connecting an *auxiliary electrode* (7.1.39) to the metal to be protected

Note 1 to entry: Galvanic protection can be cathodic or anodic.

7.4.5

impressed current protection

electrochemical protection (7.4.1) in which the protection current is supplied by an external source of electric energy

Note 1 to entry: Impressed current protection can be cathodic or anodic.

7.4.6

electrical drainage protection

electrochemical protection (7.4.1) against *stray-current corrosion* (4.14) by draining the stray-current from the metallic object

Note 1 to entry: The drainage can be obtained, for example, by connecting the metal to be protected to the negative part of the stray-current source.

7.4.7

protection potential range

range of *corrosion potential* (7.1.20) values in which an acceptable *corrosion resistance* (3.13) is achieved for a particular purpose

7.4.8

protection potential

threshold value of the *corrosion potential* (7.1.20) that has to be reached to enter a *protection potential range* (7.4.7)

7.4.9

protection current density

current density (7.2.6) that is required to maintain the *corrosion potential* (7.1.20) in a *protection potential range* (7.4.7)

7.4.10

galvanic anode

sacrificial anode

metallic component used as an *anode* (7.1.4) for *galvanic protection* (7.4.4) and *cathodic protection* (7.4.3)

Note 1 to entry: The galvanic *anode* (7.1.4) shall have a lower *corrosion potential* (7.1.20) than the metal that is to be protected.

7.4.11

insoluble anode

DEPRECATED: dimensionally stable anode

anode (7.1.4) used for *cathodic protection* (7.4.3) by *impressed current protection* (7.4.5) where the anode is not significantly consumed

EXAMPLE Platinized titanium.

7.5 Electrochemical corrosion tests

7.5.1

potentiostatic test

electrochemical test in which the *electrode potential* (7.1.18) is constantly maintained

7.5.2

potentiodynamic test

DEPRECATED: potentiokinetic test

electrochemical test in which the *electrode potential* (7.1.18) is varied continuously at a pre-set rate

7.5.3

potentiostep test

electrochemical test in which the *electrode potential* (7.1.18) is varied stepwise with time in a pre-set manner

7.5.4

potentiostaircase test

DEPRECATED: quasi-potentiostatic test

potentiostep test ([7.5.3](#)) in which the potential increments or decrements and their duration are equal for each step

7.5.5

galvanostatic test

DEPRECATED: intensiostatic test

electrochemical test in which the *current density* ([7.2.6](#)) is maintained constant

7.5.6

galvanodynamic test

electrochemical test in which the *current density* ([7.2.6](#)) is varied continuously with time at a pre-set rate

7.5.7

galvanostep test

electrochemical test in which the *current density* ([7.2.6](#)) is varied stepwise with time in a pre-set manner

7.5.8

galvanostaircase test

DEPRECATED: quasi-galvanostatic test

DEPRECATED: quasi-intensiostatic test

galvanostep test ([7.5.7](#)) in which the current increments or decrements and their duration are equal for each step

7.5.9

electrochemical impedance spectroscopy

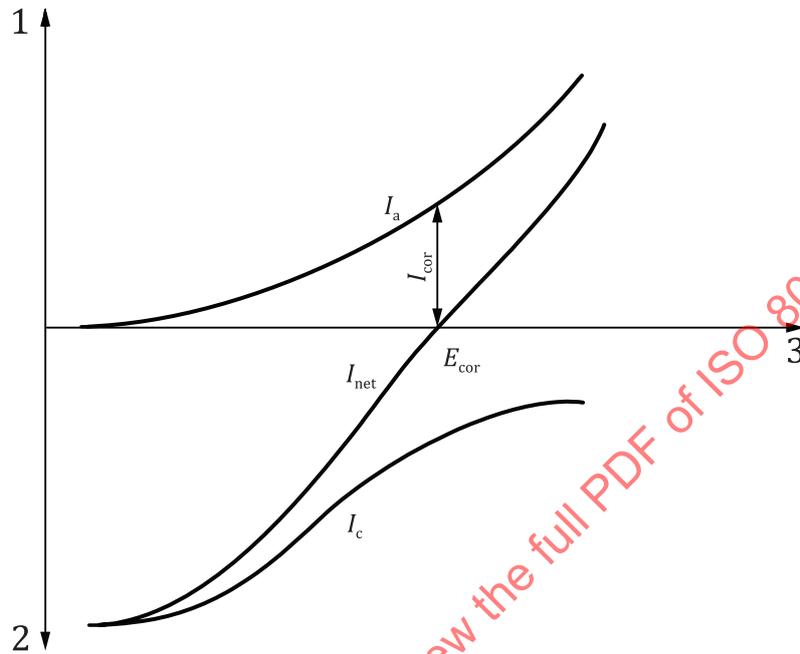
EIS

electrochemical test based on the response of a corroding *electrode* ([7.1.2](#)) to small-amplitude alternating potential or current signals at various frequencies

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Annex A
(informative)

Graphical representations of certain terms



Key

- | | | | |
|-----------|---------------------------------------------------------------------------------------------------------|-----------|--------------------------|
| 1 | anodic current | I_c | cathodic partial current |
| 2 | cathodic current | I_{net} | net current |
| 3 | electrode potential | I_{cor} | free corrosion current |
| E_{cor} | free corrosion potential | | |
| I_a | anodic partial current (in the absence of other anodic reactions, identical with the corrosion current) | | |

Figure A.1 — Current/potential curves for a corroding electrode